

NAME: _____

Chem20L, Elementary Chemistry Lab

Atomic Structure Post-Lab

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1.) Complete the following table.

Atomic Notation	# of protons (p ⁺)	# of electrons (e ⁻)	# of neutrons (n)	Ionic Charge	Atomic Number (Z)	Mass Number (A)
$^{46}_{21}\text{Sc}^{3+}$		18				
	53		72	-1		
$^{115}_{49}\text{In}$						115
$^{14}_7\text{N}^{3-}$						
				+3	74	184

2.) Identify the following elements by their chemical symbol.

a.) The alkaline earth metal in period 3.

b.) The metalloid in Group 6A.

c.) The diatomic element in Group 5A.

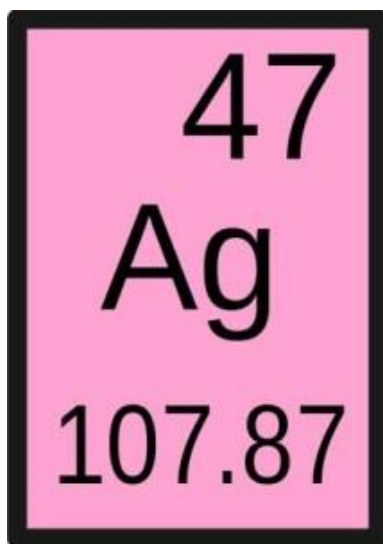
d.) The liquid at room temperature in period 4.

e.) The gas at room temperature in period 5.

3.) Write atomic notation for the following atoms.

- a.) 4 protons, 2 electrons, 5 neutrons
- b.) 76 protons, 68 electrons, 114 neutrons
- c.) 15 protons, 15 electrons, 15 neutrons
- d.) 18 protons, 18 electrons, 22 neutrons

4.) Consider the following element.



- a.) What is the element's name?
- b.) What is its atomic number?
- c.) What is its chemical symbol?
- d.) Is it a metal, nonmetal, or metalloid?
- e.) Is it a main group element or transition metal?
- f.) What is this element's atomic mass?
- g.) What period is this element in?